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Contributions of the lead-bromine weighted bands to the occupied density of states of the hybrid tri-bromide perovskites

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The electronic structure of methylammonium lead bromide ($\text{CH}_3\text{NH}_3\text{PbBr}_3$) single crystals has been investigated through a combination of resonant photoemission and theoretical modeling. There are Pb spectral contributions throughout the valence band. Importantly, the electronic structure at the top of the valence band is found to be dominated by the hybridized Pb-Br bands, not methylammonium bromide. The results line up with the partial density of states obtained from density functional theory and confirm that much of the valence band has some Pb spectral weight.

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The past few years have been witness to a series of major breakthroughs in the field of hybrid organo-lead perovskites, specifically the methyl ammonium lead trihalide perovskites MAPbX_3 ($\text{MA} = \text{CH}_3\text{NH}_3$, $\text{X} = \text{Cl}$, Br , I), as illustrated in Fig. 1. This intense attention is partly the result of better electronic properties exhibited by the single crystal materials, ranging from low defect state densities and long carrier diffusion lengths to high carrier mobilities coupled with the ease of solution processibility.¹⁻⁴ Since the first demonstration of perovskite-based photovoltaics in 2009, highly efficient solar cells have been reported with certified power conversion efficiencies (PCE) breaching 22%.⁵ Significant inroads are, as well, being made toward achieving long-term stability.⁶ Parallel efforts have been made at demonstrating light emission,⁷ photodetection⁸ and lasing applications of hybrid perovskites.⁹

There exist band structure calculations,¹⁰⁻²⁷ valence band photoemission,^{12,27-37} and combined photoemission and inverse photoemission studies,^{26,28,37} aimed at elucidating the interfacial electronic band structure of the hybrid methyl ammonium-lead halide perovskites, MAPbX_3 . Despite these attempts, however, the nature of lead (Pb) hybridization across the valence band remains poorly understood. This is important because the electronic structure at the top of the valence band plays a central role in photo-excitation and carrier transport, and an understanding of the region in the vicinity of the valence band maximum is important to the optoelectronic properties. Here, we have combined resonant

photoemission spectroscopy and density functional theory (DFT) calculations to uncover the spectral contributions to valence band electronic structure of MAPbBr_3 single crystals.

The crystals were synthesized using an antisolvent vapor-triggered crystallization approach reported elsewhere.²

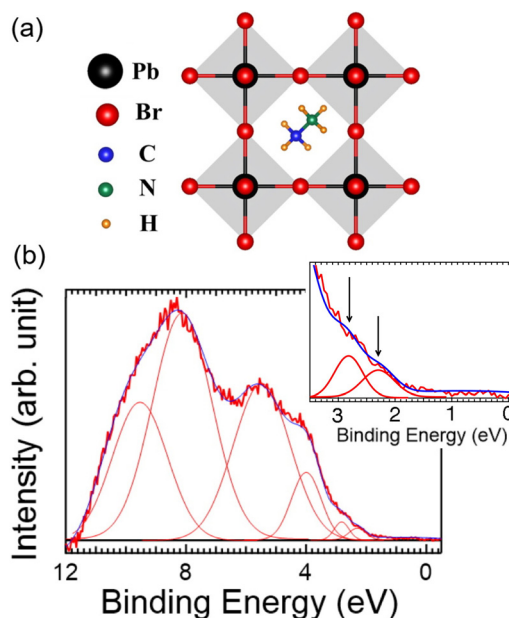


FIG. 1. (a) The cubic structure of MAPbBr_3 . (b) The valence band spectra of a MAPbBr_3 single crystal determined by synchrotron-based photoemission spectroscopy. The valence band features (at binding energies around 2.3 and 2.8 eV), in the vicinity of the valence band maximum (1.8 eV²⁶), are shown in higher detail, in the inset. Taken at a photon energy of 35 eV and binding energies, plotted as $E_F - E$.

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We have chosen the MAPbBr₃ perovskite single crystals over the other types, since the stability of these crystals has been demonstrated in the ambient,³⁸ while methylammonium lead triiodide (MAPbI₃) crystals are known to degrade in ambient conditions.³⁹ The crystals were ca. 5 mm × 3 mm × 2 mm in dimensions.

To obtain detailed information on the states at the top of the valence band, we employed synchrotron-based photon energy-dependent photoemission spectroscopy. The studies were carried out at the photoemission end station of the plane grating mirror (PGM) beamline at Canadian Light Source, Inc., equipped with a Scienta SES100 hemispherical electron energy analyzer, with a total energy resolution of 50 meV. All the spectra were collected at normal emission and room temperature from MAPbBr₃(001) single crystals. Photon energies were tunable in the range of a few eV to a few-hundred eV. Spectra were calibrated using the Fermi edge of a sputter-cleaned Cu sample. The spot size of the photon beam was on the order of microns.

Figure 1(a) shows the typical crystal structure of the cubic MAPbBr₃, along the (001) surface normal. Figure 1(b) shows the photoemission spectrum acquired from a MAPbBr₃(001) single crystal. Four main valence band features were observed at the binding energies ($E_F - E$) of roughly 4 eV, 5.4 eV, 8 eV, and 9.6 eV. Two weak intensity features in the valence band photoemission spectra are located at binding energies of 2.3 and 2.8 eV, as indicated by arrows, and the fittings to the experimental data, in the inset of Fig. 1(b). These states are close, in binding energy, to the top of the valence band which is seen to be at a binding energy of 1.8 eV, as measured at the Brillouin zone edge in a recent experimental band structure.²⁶ Since the surface termination of MAPbBr₃(001) is MABr,²⁵ these states observed in photoemission are consistent with bulk band states of the associated calculated band structure.^{25,26} Experimental verification of the bulk band structure character at the top of the valence band would require mapping the photon energy dependence of the valence band at the surface Brillouin zone edge high symmetry point \bar{M} ,²⁶ which was not done here. Such measurements would establish the presence or absence of dispersion in k_{\perp} (the presence or absence of conservation of 2-dimensionality of state), although it is clear that the top of the valence at \bar{M} does not fall into a gap of the projected bulk band density of states.²⁶ Such dispersion, in k_{\perp} at the surface Brillouin zone edge, but with wave vector along the surface normal, i.e., a photon energy dependent study, using angle resolved photoemission, was performed for methylammonium lead iodide (CH₃NH₃PbI₃) single crystals,³² establishing the bulk band character of the top of the valence band. For peak fitting, we used a Shirley background, with subsequent background subtraction. The valence band spectral features [Fig. 1(b)] were fitted with Gaussians as follows: a peak at a binding energy of 2.3 eV ($E_F - E$) with a full width at half maximum (FWHM) of 0.66 eV; a peak at 2.8 eV with a FWHM 0.58 eV; a peak at 3.9 eV with a FWHM 1.16 eV; a peak at 5.5 eV with a FWHM 2.25 eV; a peak at 8.1 eV with a FWHM 2.33 eV; and a peak at 10 eV with a FWHM 2.16 eV.

To experimentally determine the lead spectral weight contributions to the valence band, especially near the top of

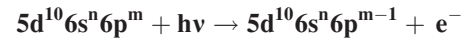
the valence band, we employed an approach based on resonant photoemission enhancement of the *partially-occupied* valence Pb 6p levels. Resonant photoemission involves the resonance photoemission enhancement of the valence band by one of the shallow core levels of the system under study,^{40–43} and in this case the resonant process is



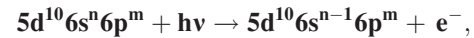
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where an electron is excited to an unoccupied conduction band state, of 6p character, which then collapses, causing resonant enhancement for occupied states containing either Pb 6p or Pb 6s spectral weight density. This leads to resonant enhancement of the direct photoemission process



or



respectively, and results in the same final state.⁴³ The resonance enhancement in photoemission yields direct insights into elemental spectral contributions to a specific valence state and/or top-most valence band features of the material.

A 2D map of the resonant photoemission intensity for the valence band spectral features, in the binding energy region 0–5 eV, was acquired as a function of the photon energy ($h\nu = 14$ –35 eV) and is shown in Fig. 2(a). The partial density of states (PDOS) intensity profile of the spectral features located around *ca.* 2.25 eV binding energy [indicated by vertical blue line in Fig. 2(a)] reveals the resonant enhancement at the photon energies of 22 eV and 24.5 eV [Fig. 2(b)]. These photon energies for the peak resonant photoemission occur at the Pb 5d_{5/2} and 5d_{3/2} shallow core level thresholds at binding energies of 19.8 eV and 22.4 eV, respectively, as plotted in Fig. 3, plus the band gap of 2.3 eV,^{11,26,44–48} as expected. Thus, we attribute the resonant enhancement to photoemission resonant transition from Pb 5d to a partially unoccupied Pb 6p states, in the conduction band, and indicates a Pb contribution to most of the valence band, including the region near the top of the valence band. This type of resonant photoemission enhancement, of almost the entire valence band, has been seen for other perovskites, including the manganese perovskites.^{40–43} The very strong resonant enhancement seen here, in the photoemission as a function of photon energy, is also characteristic of a nonmetallic system,⁴³ and indeed is expected as MAPbBr₃ is a semiconductor, although metallic lead surface segregation can be a problem.^{26,30,31,35} Several theoretical^{10–27} and experimental^{26,32} papers have the dispersion near the top of the valence band to be in the region of 1 eV or less, and the dispersion of the heavy mass Pb 5p-Br 4p bands (at a binding energy of 4 eV) to be in the region of several hundred meV or less. Thus, the results of Fig. 2(a) are not the result of bulk band dispersion, with changing photon energy.

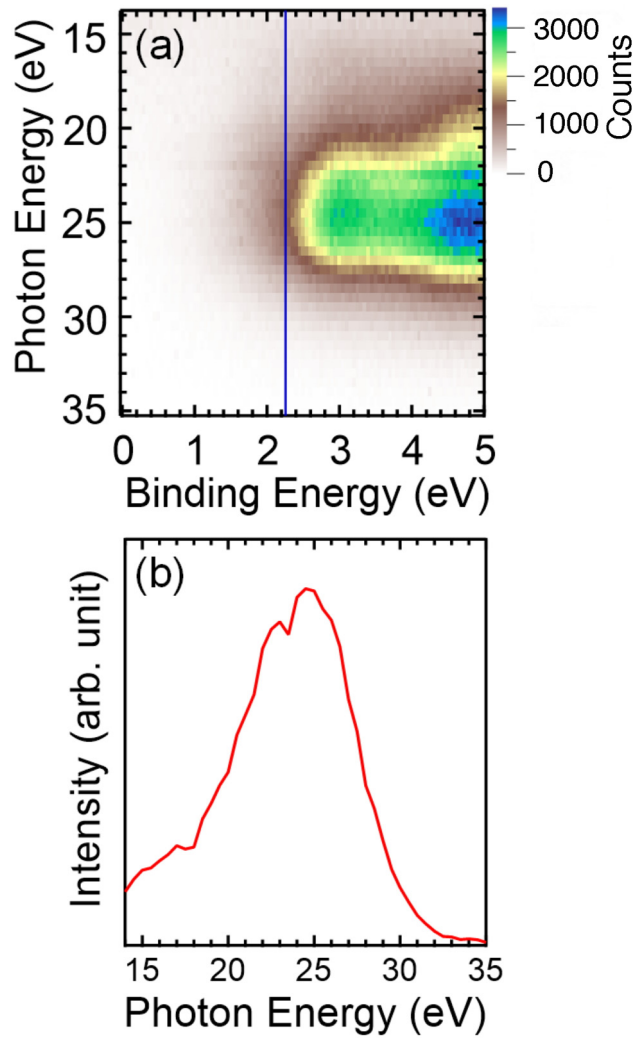


FIG. 2. (a) The results from resonant photoemission, constant initial state spectra, illustrated as a 2D map of the valence band photoemission intensities, as a function of the photon energy. (b) Resonant enhancement of the valence band in the region 2.3 eV binding energy ($E_F - E$) as a function of photon energy, with the greatest enhancements occurring at photon energies corresponding to the binding energies of the Pb 5d shallow core levels.

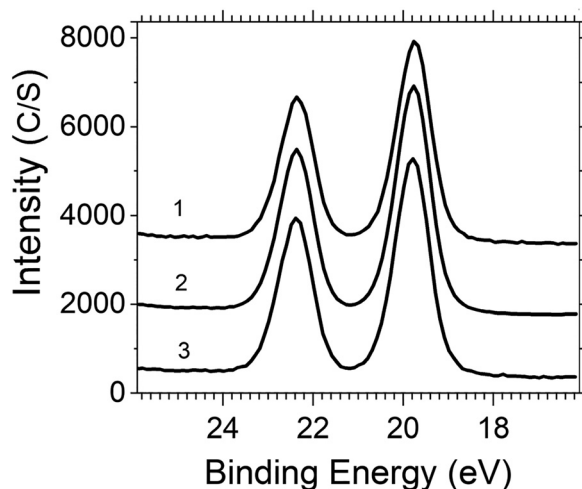


FIG. 3. The Pb 5d core levels from three different MAPbBr₃ single crystals, as established by X-ray photoemission using a PHI Versa Probe II instrument equipped with monochromatic Al K(α) source. Spectra are offset for clarity.

To further confirm the spectral contributions to the states at the top of the valence band, we calculated the electronic band structure of MAPbBr₃ using density functional theory (DFT). We employed the projected augmented wave (PAW) method for the electron-ion potential,⁴⁹ the generalized gradient approximation (GGA) for exchange-correlation potential,⁵⁰ and relativistic corrections including spin-orbit coupling as implemented in Vienna *ab-initio* simulation package (VASP).^{51,52} The calculations were carried out using a 640 eV kinetic energy cutoff and $12 \times 12 \times 12$ k-point mesh. We fully relaxed the ionic coordinates with the force convergence limit of 0.001 eV/atom. The theoretical lattice constant of 5.96 Å was used, and we maintained cubic lattice symmetry in the calculations. In GGA, it is well known that band gap is underestimated, and as a result we have shifted the conduction band by 1.37 eV to better match the experimental band gap of 2.33 eV.⁵³

The calculated partial density of states (PDOS) of a cubic MAPbBr₃ crystal [Fig. 1(a)] has been plotted in Fig. 4. We have used the tetrahedral method to calculate the partial density of states; therefore, Fig. 4 has no smearing included. The majority of the valence band is dominated by Br-p orbitals while the conduction band is dominated by Pb-p orbitals.²⁶ The top of the valence band, however, has almost equivalent contributions from Pb-s and Br-p orbitals,^{26,27} as applicable to a single Pb-Br layer. The MA-Br layer should dominate the photoemission as this would be the surface

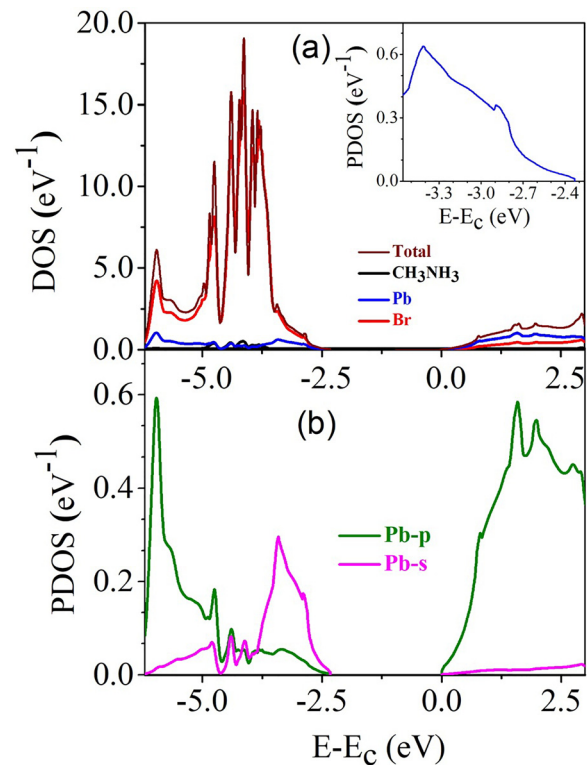


FIG. 4. Partial density of states (PDOS) of MAPbBr₃ showing (a) the contributions of Pb and Br atoms and the CH₃NH₃ molecule, in the presence of spin-orbit coupling, as well as the total density of states corrected for a full unit cell and (b) the relative contributions of Pb-p (green) and Pb-s (violet). The binding energy (E) is defined with respect to the energy of the conduction band minima (E_c). Inset shows the Pb weight at the valence band maximum.

terminal layer,^{25,26} as noted earlier. Lindblad and coworkers³⁶ have the top of the valence band dominated almost exclusively by Br spectral weight, which results from the fairly heavy overall Br spectral weight contributions, but does not correct for the surface sensitivity of photoemission. The occupied states corresponding to the CH₃NH₃ molecule lie far below the valence band maximum. Due to the large contribution of Pb-p orbitals in the conduction band, spin-orbit coupling mainly affects the conduction band. There are two states that appear in the calculated PDOS, ca. 0.5 and 1 eV below the valence band maximum, or 2.8 eV and 3.4 eV below the conduction band edge. If the valence band maximum is placed at a binding energy of 1.8 eV, as indicated by prior work on MAPbBr₃,²⁶ then in experiment, these features near the top of the valence band in the PDOS (inset of Fig. 4) would appear at 2.3 eV and 2.8 eV. This is very close to what is observed in photoemission (2.3 eV and 2.8 eV), as seen in Fig. 2.

In summary, we have explored the electronic structure of MAPbBr₃ single crystals in the vicinity of the valence band maximum, using a combination of resonant photoemission spectroscopy and density functional theory. Our studies suggest that the top of the valence band of these hybrid perovskites has spectral contributions from both Pb and Br and is dominated by Pb-Br hybridized bands. Based on other studies,^{26,32} these bands are not surface state bands, rather contain strong bulk character.

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¹G. Xing, N. Mathews, S. Sun, S. S. Lim, Y. M. Lam, M. Grätzel, S. Mhaisalkar, and T. C. Sum, *Science* **342**, 344 (2013).

²D. Shi, V. Adinolfi, R. Comin, M. Yuan, E. Alarousu, A. Buin, Y. Chen, S. Hoogland, A. Rothenberger, K. Katsiev, Y. Losovyj, X. Zhang, P. A. Dowben, O. F. Mohammed, E. H. Sargent, and O. M. Bakr, *Science* **347**, 519 (2015).

³M. I. Saidaminov, A. L. Abdelhady, B. Murali, E. Alarousu, V. M. Burlakov, W. Peng, I. Dursun, L. Wang, Y. He, G. Maculan, A. Goriely, T. Wu, O. F. Mohammed, and O. M. Bakr, *Nat. Commun.* **6**, 7586 (2015).

⁴Q. Dong, Y. Fang, Y. Shao, P. Mulligan, J. Qiu, L. Cao, and J. Huang, *Science* **347**, 967–970 (2015).

⁵M. A. Green, K. Emery, Y. Hishikawa, W. Warta, and E. D. Dunlop, *Prog. Photovoltaics Res. Appl.* **23**, 1–9 (2015).

⁶G. Grancini, C. Roldán-Carmona, I. Zimmermann, E. Mosconi, X. Lee, D. Martineau, S. Narbey, F. Oswald, F. De Angelis, M. Graetzel, and M. K. Nazeeruddin, *Nat. Commun.* **8**, 15684 (2017).

⁷H. Cho, S.-H. Jeong, M.-H. Park, Y.-H. Kim, C. Wolf, C.-L. Lee, J. H. Heo, A. Sadhanala, N. Myoung, S. Yoo, S. H. Im, R. H. Friend, and T.-W. Lee, *Science* **350**, 1222 (2015).

⁸M. I. Saidaminov, V. Adinolfi, R. Comin, A. L. Abdelhady, W. Peng, I. Dursun, M. Yuan, S. Hoogland, E. H. Sargent, and O. M. Bakr, *Nat. Commun.* **6**, 8724 (2015).

⁹H. Zhu, Y. Fu, F. Meng, X. Wu, Z. Gong, Q. Ding, M. V. Gustafsson, M. T. Trinh, S. Jin, and X. Y. Zhu, *Nat. Mater.* **14**, 636 (2015).

- ¹⁰Q. Chen, N. De Marco, Y. Yang, T.-B. Song, C.-C. Chen, H. Zhao, Z. Hong, H. Zhou, and Y. Yang, *Nanotoday* **10**, 355–396 (2015).
- ¹¹T. Shi, W.-J. Yin, F. Hong, K. Zhu, and Y. Yan, *Appl. Phys. Lett.* **106**, 103902 (2015).
- ¹²H.-J. Feng, T. R. Paudel, E. Y. Tsymlal, and X. C. Zeng, *J. Am. Chem. Soc.* **137**, 8227 (2015).
- ¹³J.-S. Park, S. Choi, Y. Yan, Y. Yang, J. M. Luther, S.-H. Wei, P. Parilla, and K. Zhu, *J. Phys. Chem. Lett.* **6**, 4304 (2015).
- ¹⁴T. Uemebayashi, K. Asai, T. Kondo, and A. Nakao, *Phys. Rev. B* **67**, 155405 (2003).
- ¹⁵W. Geng, C.-T. Tong, Z.-K. Tang, C. Yam, Y.-N. Zhang, W.-M. Lau, and L.-M. Liu, *J. Mater. Chem. C* **1**, 213–220 (2015).
- ¹⁶E. Mosconi, A. Amat, M. K. Nazeeruddin, M. Grätzel, and F. De Angelis, *J. Phys. Chem. C* **117**, 13902 (2013).
- ¹⁷T. Baikie, Y. Fang, J. M. Kadro, M. Schreyer, F. Wei, S. G. Mhaisalkar, M. Graetzel, and T. J. White, *J. Mater. Chem. A* **1**, 5628–5641 (2013).
- ¹⁸J. Even, L. Pedesseau, C. Katan, M. Kepenekian, J.-S. Lauret, D. Saporì, and E. Deleporte, *J. Phys. Chem. C* **119**, 10161–10177 (2015).
- ¹⁹J. Even, L. Pedesseau, and C. Katan, *J. Phys. Chem. C* **118**, 11566–11572 (2014).
- ²⁰J. Even, L. Pedesseau, J.-M. Jancu, and C. Katan, *J. Phys. Chem. Lett.* **4**, 2999–3005 (2013).
- ²¹M. Kepenekian, R. Robles, C. Katan, D. Saporì, L. Pedesseau, and J. Even, *ACS Nano* **9**, 11557–11567 (2015).
- ²²G. Giorgi, J.-I. Fujisawa, H. Segawa, and K. Yamashita, *J. Phys. Chem. Lett.* **4**, 4213–4216 (2013).
- ²³P. Umari, E. Mosconi, and F. De Angelis, *Sci. Rep.* **4**, 4467 (2014).
- ²⁴F. Brivio, K. T. Butler, A. Walsh, and M. van Schilfgaarde, *Phys. Rev. B* **89**, 155204 (2014).
- ²⁵X. Huang, T. R. Paudel, P. A. Dowben, S. Dong, and E. Y. Tsymlal, *Phys. Rev. B* **94**, 195309 (2016).
- ²⁶T. Komesu, X. Huang, T. R. Paudel, Y. B. Losovyj, X. Zhang, E. F. Schwier, Y. Kojima, M. Zheng, H. Iwasawa, K. Shimada, M. I. Saidaminov, D. Shi, A. L. Abdelhady, O. M. Bakr, S. Dong, E. Y. Tsymlal, and P. A. Dowben, *J. Phys. Chem. C* **120**, 21710 (2016).
- ²⁷J. Qian, B. Xu, and W. Tian, *Org. Electron.* **37**, 61–73 (2016).
- ²⁸C. Wang, X. Liu, C. Wang, Z. Xiao, C. Bi, Y. Shao, J. Huang, and Y. Gao, *J. Vac. Sci. Technol., B* **33**, 032401 (2015).
- ²⁹M. Kollár, L. Čirić, J. H. Dil, A. Weber, S. Muff, H. M. Ronnow, B. Náfrádi, B. P. Monnier, J. S. Luterbacher, L. Forró, and E. Horváth, *Sci. Rep.* **7**, 695 (2017).
- ³⁰G. Sadoughi, D. E. Starr, E. Handick, S. D. Stranks, M. Gorgoi, R. G. Wilks, M. Bär, and H. J. Snaith, *ACS Appl. Mater. Interfaces* **7**, 13440–13444 (2015).
- ³¹B. Conings, L. Baeten, C. De Dobbelaere, J. D'Haen, J. Manca, and H.-G. Boyen, *Adv. Mater.* **26**, 2041–2046 (2014).
- ³²M.-I. Lee, A. Barragán, M. N. Nair, V. L. R. Jacques, D. Le Bolloc'h, P. Fertey, K. Jemli, F. Lédée, G. Trippé-Allard, E. Deleporte, A. Taleb-Ibrahimi, and A. Tejada, *J. Phys. D: Appl. Phys.* **50**, 26LT02 (2017).
- ³³R. Lindblad, D. Bi, B.-W. Park, J. Oscarsson, M. Gorgoi, H. Siegbahn, M. Odelius, E. M. J. Johansson, and H. Rensmo, *J. Phys. Chem. Lett.* **5**, 648–653 (2014).
- ³⁴D. Niesner, M. Wilhelm, I. Levchuk, A. Osvet, S. Shrestha, M. Batentschuk, C. Brabec, and T. Fauster, *Phys. Rev. Lett.* **117**, 126401 (2016).
- ³⁵F.-S. Zu, P. Amsalem, I. Salzmann, R.-B. Wang, M. Ralaivisoa, S. Kowarik, S. Duhm, and N. Koch, *Adv. Opt. Mater.* **5**, 1700139 (2017).
- ³⁶R. Lindblad, N. K. Jena, B. Philippe, J. Oscarsson, D. Bi, A. Lindblad, S. Mandal, B. Pal, D. D. Sarma, O. Karis, H. Siegbahn, E. M. J. Johansson, M. Odelius, and H. Rensmo, *J. Phys. Chem. C* **119**, 1818–1825 (2015).
- ³⁷J. Endres, D. A. Egger, M. Kulbak, R. A. Kerner, L. Zhao, S. H. Silver, G. Hodes, B. P. Rand, D. Cahen, L. Kronik, and A. Kahn, *J. Phys. Chem. Lett.* **7**, 2722–2729 (2016).
- ³⁸H.-H. Fang, S. Adjokatse, H. Wei, J. Yang, G. R. Blake, J. Huang, J. Even, and M. A. Loi, *Sci. Adv.* **2**, e1600534 (2016).
- ³⁹J. F. Galisteo-López, M. Anaya, M. E. Calvo, and H. Míguez, *J. Phys. Chem. Lett.* **6**, 2200 (2015).
- ⁴⁰D. N. McIlroy, J. Zhang, S.-H. Liou, and P. A. Dowben, *Phys. Lett. A* **207**, 367–373 (1995).
- ⁴¹J. Zhang, D. N. McIlroy, P. A. Dowben, S.-H. Liou, S. F. Sabiryanov, and S. S. Jaswal, *Solid State Commun.* **97**, 39–44 (1996).
- ⁴²D. N. McIlroy, C. Waldfried, J. Zhang, J.-W. Choi, F. Foong, S.-H. Liou, and P. A. Dowben, *Phys. Rev. B* **54**, 17438–17451 (1996).
- ⁴³P. A. Dowben, *Surf. Sci. Rep.* **40**, 151 (2000).
- ⁴⁴R. Sheng, A. Ho-Baillie, S. Huang, S. Chen, X. Wen, X. Hao, and M. A. Green, *J. Phys. Chem. C* **119**, 3545–3549 (2015).

- ⁴⁵Y. Yang, M. Yang, Z. Li, R. Crisp, K. Zhu, and M. C. Beard, *J. Phys. Chem. Lett.* **6**, 4688–4692 (2015).
- ⁴⁶D. Priante, I. Dursun, M. S. Alias, D. Shi, V. A. Melnikov, T. K. Ng, O. F. Mohammed, O. M. Bakr, and B. S. Ooi, *Appl. Phys. Lett.* **106**, 081902 (2015).
- ⁴⁷N. Kitazawa, Y. Watanabe, and Y. Nakamura, *J. Mater. Sci.* **37**, 3585 (2002).
- ⁴⁸K. Tanaka, T. Takahashi, T. Ban, T. Kondo, K. Uchida, and N. Miura, *Solid State Commun.* **127**, 619–623 (2003).
- ⁴⁹P. E. Blöchl, *Phys. Rev. B* **50**, 17953 (1994).
- ⁵⁰J. P. Perdew, K. Burke, and M. Ernzerhof, *Phys. Rev. Lett.* **77**, 3865 (1996).
- ⁵¹G. Kresse and J. Furthmüller, *Phys. Rev. B* **54**, 11169 (1996).
- ⁵²G. Kresse and D. Joubert, *Phys. Rev. B* **59**, 1758 (1999).
- ⁵³G. C. Papavassiliou and I. B. Koutselas, *Synth. Met.* **71**, 1713 (1995).